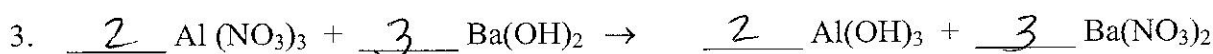


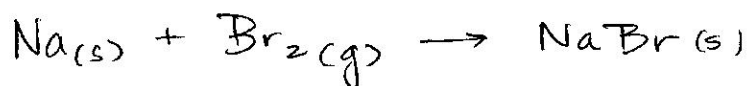
I. Balance the following equations:

Is equation # 4 an endothermic or an exothermic reaction? (circle one)

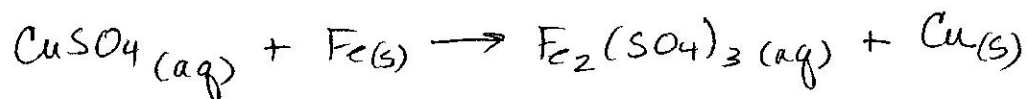
List the reactants in reaction #1 H_2S and Cl_2

II. Write chemical equations from the following word equations. *Be sure to note the physical state of each reactant and product*****

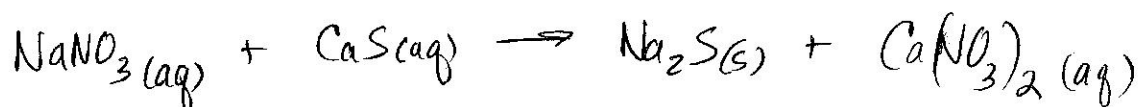
1. Sodium metal combines with bromine gas to form sodium bromide



2. Aqueous copper (II) sulfate reacts with solid Iron to produce aqueous Iron (III) sulfate and solid copper.



3. Aqueous sodium nitrate combines with aqueous calcium sulfide to form solid sodium sulfide and aqueous calcium nitrate.



III. Identify the physical state of each of the following compounds as solid (s) or aqueous (aq)

S AgCl

S PbSO₄

aq K₂S

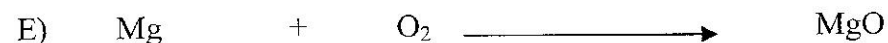
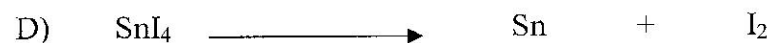
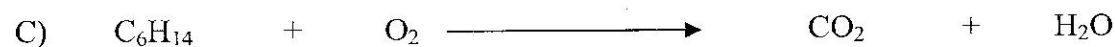
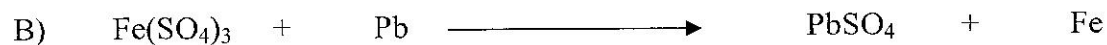
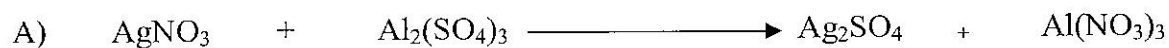
aq (NH₄)₂S

aq Aluminum Nitrate

S Calcium Carbonate

aq Magnesium Iodide

IV. Determine the type of reaction.



V. Name the types of equations in A-E

A) Double Displacement

B) Single Displacement

C) Combustion

D) Decomposition

E) Synthesis